

## Comment

### Affinity dye–ligand poly(hydroxyethyl methacrylate)/chitosan composite membrane for adsorption lysozyme and kinetic properties G. Bayramoğlu, M. Yilmaz, M.Y. Arica

Recently, Bayramoğlu et al. [1] published the paper entitled as above. In Section 2, pseudo-first and second order equations, Bayramoğlu et al. mentioned that a pseudo-second order equation based on adsorption equili-

brium capacity may be expressed in the form:

$$\frac{dq_t}{dt} = k_2(q_{eq} - q_t)^2$$

In fact it was Ho [2], who first developed a pseudo-second order kinetic expression for the sorption systems of divalent metal ions using sphagnum moss peat. In earlier application of the pseudo-second order equation to the kinetic studies of competitive heavy metal adsorption by sphagnum moss peat was undertaken by Ho et al. [3]. The pseudo-second order rate expression of Ho has been applied to the sorption of metal ions, dyes and organic substances from aqueous solution (Table 1). In addition, Ho's kinetic expression has also been applied to the sorption of dye onto the mixture sorbent sorption processes [4]. Numerous applications of Ho's kinetic expression have been reported in recent years. A list of pseudo-second order systems is given in Table 1.

I suggest that Bayramoğlu et al. should cite Ho's original pseudo-second order kinetic expression paper.

Table 1  
Pseudo-second order kinetic model of various related systems from the literature

Sorbent	Sorbate	Reference
<i>C. vulgaris</i>	Cd(II)	[5]
<i>Rhizopus arrhizus</i>	Remazol Black B	[6]
<i>Trametes versicolor</i>	Cd(II)	[7]
<i>Arundo canes</i>	Cd(II), Ni(II)	[8]
Zeolite tuff	Pb(II)	[9]
Algae	Lauryl benzyl sulfonate	[10]
<i>Aspergillus niger</i>	Basic Blue 9	[11]
<i>Aspergillus niger</i>	Acid Blue 29	[12]
Mixed clay/carbon	Acid Blue 9	[4]
Wood	Basic Blue 69, Acid Blue 25	[13]
Peat	Basic Blue 69, Acid Blue 25	[14]
Sphagnum moss peat	Chrysoidine (BO2), Astrazon Blue (BB3), Astrazone Blue (BB69)	[15]
Pith	Basic Red 22, Acid Red 114	[16]
Fly ash	Omega Chrome Red ME, <i>o</i> -cresol, <i>p</i> -nitrophenol	[17]
Sphagnum moss peat	Cu(II), Ni(II)	[3]
Sphagnum moss peat	Cu(II), Ni(II), Pb(II)	[18]
Activated clay	Basic Red 18, Acid Blue 9	[19]
<i>Aspergillus niger</i>	Pb(II), Cd(II), Cu(II), Ni(II)	[20]
Activated carbon	Hg(II)	[21]
Spent grain	Pb(II), Cd(II)	[22]
2-Mercaptobenzimidazole-clay	Hg(II)	[23]
Perlite	Cd(II)	[24]
Coir	Cu(II), Pb(II)	[25]
Sago	Cu(II), Pb(II)	[26]
Glass	Cu(II)	[27]
Polysaccharide	Pb(II), Cu(II), Zn(II), Cd(II), Ni(II)	[28]
Chitin, Chitosan, <i>Rhizopus arrhizus</i>	Cr(VI), Cu(II)	[29]
Microcystis	Ni(II), Cr(VI)	[30]
<i>Aspergillus niger</i>	Congo Red	[31]
<i>Rhizopus arrhizus</i>	Fe(III), Fe(III)-cyanide complex	[32]
Banana stalk ( <i>Musa paradisiaca</i> )	Hg(II)	[33]

## References

- [1] G. Bayramoğlu, M. Yilmaz, M.Y. Arica, Affinity dye–ligand poly(hydroxyethyl methacrylate)/chitosan composite membrane for adsorption lysozyme and kinetic properties, *Biochem. Eng. J.* 13 (2003) 35–42.
- [2] Y.S. Ho, Adsorption of Heavy Metals from Waste Streams by Peat, Ph.D. Thesis, University of Birmingham, Birmingham, UK, 1995.
- [3] Y.S. Ho, D.A.J. Wase, C.F. Forster, Kinetic studies of competitive heavy metal adsorption by sphagnum moss peat, *Environ. Technol.* 17 (1996) 71–77.
- [4] Y.S. Ho, C.C. Chiang, Sorption studies of acid dye by mixed sorbents, *Adsorpt. J. Int. Adsorpt. Soc.* 7 (2001) 139–147.
- [5] Z. Aksu, Equilibrium and kinetic modelling of cadmium(II) biosorption by *C. Vulgaris* in a batch system: effect of temperature, *Sep. Purif. Technol.* 21 (2001) 285–294.
- [6] Z. Aksu, S. Tezer, Equilibrium and kinetic modelling of biosorption of Remazol Black B by *Rhizopus arrhizus* in a batch system: effect of temperature, *Process Biochem.* 36 (2000) 431–439.
- [7] M.Y. Arica, Y. Kaçar, Ö. Genç, Entrapment of white-rot fungus *Trametes versicolor* in Ca-alginate beads: preparation and biosorption kinetic analysis for cadmium removal from an aqueous solution, *Bioresour. Technol.* 80 (2001) 121–129.
- [8] M.C. Basso, E.G. Cerrella, A.L. Cukierman, Activated carbons developed from a rapidly renewable biosource for removal of cadmium(II) and nickel(II) ions from dilute aqueous solutions, *Ind. Eng. Chem. Res.* 41 (2002) 180–189.
- [9] R.F. El-Bishtawi, A.A.H. Ali, Sorption kinetics of lead ions by zeolite tuff, *J. Environ. Sci. Health Part A. Toxic/Hazard. Subst. Environ. Eng.* 36 (2001) 1055–1072.
- [10] N.A. Fernandez, E. Chacin, E. Gutierrez, N. Alastre, B. Llamaza, C.F. Forster, Adsorption of lauryl benzyl sulfonate on algae, *Bioresour. Technol.* 54 (1995) 111–115.

- [11] Y.Z. Fu, T. Viraraghavan, Removal of a dye from an aqueous solution by the fungus *Aspergillus niger*, *Water Qual. Res. J. Can.* 35 (2000) 95–111.
- [12] Y.Z. Fu, T. Viraraghavan, Removal of CI Acid Blue 29 from an aqueous solution by *Aspergillus niger*, *AATCC Rev.* 1 (2001) 36–40.
- [13] Y.S. Ho, G. McKay, Kinetic models for the sorption of dye from aqueous solution by wood, *Process Saf. Environ. Protect.* 76 (B2) (1998) 183–191.
- [14] Y.S. Ho, G. McKay, Sorption of dye from aqueous solution by peat, *Chem. Eng. J.* 70 (1998) 115–124.
- [15] Y.S. Ho, G. McKay, The kinetics of sorption of basic dyes from aqueous solution by sphagnum moss peat, *Can. J. Chem. Eng.* 76 (1998) 822–827.
- [16] Y.S. Ho, G. McKay, A kinetic study of dye sorption by biosorbent waste product pith, *Resour. Conserv. Recycl.* 25 (1999) 171–193.
- [17] Y.S. Ho, G. McKay, Comparative sorption kinetic studies of dye and aromatic compounds onto fly ash, *J. Environ. Sci. Health Part A. Toxic/Hazard. Subst. Environ. Eng.* 34 (1999) 1179–1204.
- [18] Y.S. Ho, G. McKay, The kinetics of sorption of divalent metal ions onto sphagnum moss peat, *Water Res.* 34 (2000) 735–742.
- [19] Y.S. Ho, C.C. Chiang, Y.C. Hsu, Sorption kinetics for dye removal from aqueous solution using activated clay, *Sep. Sci. Technol.* 36 (2001) 2473–2488.
- [20] A. Kapoor, T. Viraraghavan, D.R. Cullimore, Removal of heavy metals using the fungus *Aspergillus niger*, *Bioresour. Technol.* 70 (1999) 95–104.
- [21] K.A. Krishnan, T.S. Anirudhan, Removal of mercury(II) from aqueous solutions and chlor-alkali industry effluent by steam activated and sulphurised activated carbons prepared from bagasse pith: kinetics and equilibrium studies, *J. Hazard. Mater.* 92 (2002) 161–183.
- [22] K.S. Low, C.K. Lee, S.C. Liew, Sorption of cadmium and lead from aqueous solutions by spent grain, *Process Biochem.* 36 (2000) 59–64.
- [23] D.M. Manohar, K.A. Krishnan, T.S. Anirudhan, Removal of mercury(II) from aqueous solutions and chlor-alkali industry wastewater using 2-mercaptobenzimidazole-clay, *Water Res.* 36 (2002) 1609–1619.
- [24] T. Mathialagan, T. Viraraghavan, Adsorption of cadmium from aqueous solutions by perlite, *J. Hazard. Mater.* 94 (2002) 291–303.
- [25] S.Y. Quek, B. Al Duri, D.A.J. Wase, C.F. Forster, Coir as a biosorbent of copper and lead, *Process Saf. Environ. Protect.* 76 (B1) (1998) 50–54.
- [26] S.Y. Quek, D.A.J. Wase, C.F. Forster, The use of sago waste for the sorption of lead and copper, *Water SA* 24 (1998) 251–256.
- [27] B.J. Rappoli, D.A. Rowley, The sorption kinetics of copper(II) on chemically modified controlled pore glass, *J. Colloid Interf. Sci.* 226 (2000) 218–221.
- [28] Z. Reddad, C. Gérente, Y. Andres, P. Le Cloirec, Adsorption of several metal ions onto a low-cost biosorbent: kinetic and equilibrium studies, *Environ. Sci. Technol.* 36 (2002) 2067–2073.
- [29] Y. Sağ, Y. Aktay, Kinetic studies on sorption of Cr(VI) and Cu(II) ions by chitin, chitosan and *Rhizopus arrhizus*, *Biochem. Eng. J.* 12 (2001) 143–153.
- [30] S. Singh, B.N. Rai, L.C. Rai, E. Pişkin, Ni(II) and Cr(VI) sorption kinetics by microcystis in single and multimetallic system, *Process Biochem.* 36 (2001) 1205–1213.
- [31] Y. Fu, T. Viraraghavan, Removal of Congo Red from an aqueous solution by fungus *Aspergillus niger*, *Adv. Environ. Res.* 7 (2002) 239–247.
- [32] Z. Aksu, H. Gülen, Binary biosorption of iron(III) and iron(III)-cyanide complex ions on *Rhizopus arrhizus*: modelling of synergistic interaction, *Process Biochem.* 38 (2002) 161–173.
- [33] I.G. Shibi, T.S. Anirudhan, Synthesis, characterization, and application as a mercury(II) sorbent of banana stalk (*Musa paradisiaca*)—polyacrylamide grafted copolymer bearing carboxyl groups, *Ind. Eng. Chem. Res.* 41 (2002) 5341–5352.

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